

This Page Is Inserted by IFW Operations  
and is not a part of the Official Record

## **BEST AVAILABLE IMAGES**

Defective images within this document are accurate representations of the original documents submitted by the applicant.

Defects in the images may include (but are not limited to):

- BLACK BORDERS
- TEXT CUT OFF AT TOP, BOTTOM OR SIDES
- FADED TEXT
- ILLEGIBLE TEXT
- SKEWED/SLANTED IMAGES
- COLORED PHOTOS
- BLACK OR VERY BLACK AND WHITE DARK PHOTOS
- GRAY SCALE DOCUMENTS

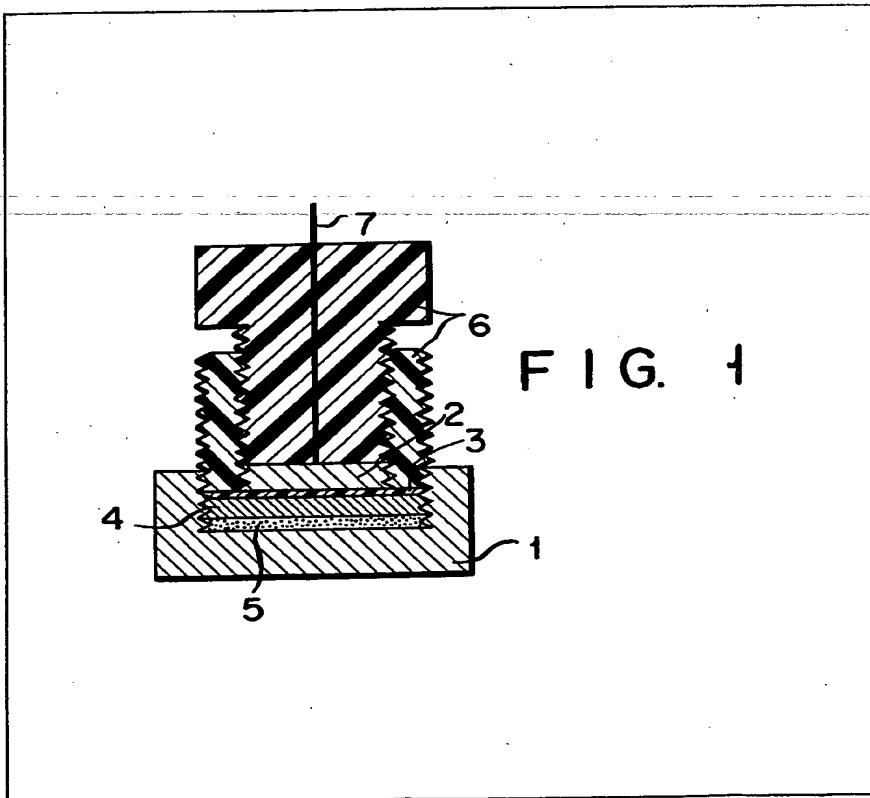
## **IMAGES ARE BEST AVAILABLE COPY.**

**As rescanning documents *will not* correct images,  
please do not report the images to the  
Image Problem Mailbox.**

(21) Application No 8039292  
 (22) Date of filing 8 Dec 1980  
 (30) Priority data  
 (31) 54/163622  
 54/163621  
 54/163623  
 55/003801  
 55/021575  
 (32) 18 Dec 1979  
 18 Dec 1979  
 18 Dec 1979  
 17 Jan 1980  
 25 Feb 1980  
 (33) Japan (JP)  
 (43) Application published  
 12 Aug 1981  
 (51) INT CL<sup>3</sup>  
 H01M 4/60  
 (52) Domestic classification  
 H1B 460 462 614A 616 618  
 620  
 (56) Documents cited  
 GB 2040548A  
 GB 2023329A  
 GB 1542325  
 GB 1517745  
 GB 1358999  
 GB 1352648  
 GB 1237044  
 GB 1216549  
 (58) Field of search  
 H1B  
 (71) Applicants  
 Nippon Telegraph &  
 Telephone Public  
 Corporation,  
 1-6, 1-chome,  
 Uchisaiwai-cho,  
 Chiyoda-ku,  
 Tokyo,  
 Japan.  
 (72) Inventors  
 Shinichi Tobishima,  
 Junichi Yamaki,  
 Akihiko Yamaji.  
 (74) Agents  
 A.A. Thornton & Co.,  
 Northumberland House,  
 303-306 High Holborn,  
 London WC1V 7LE.

## (54) Light weight battery

(57) The battery comprises an anode (2) including an anode-active material formed of one metal selected from the Group Ia metals or preferably lithium metal, a cathode (5) including a cathode-active material formed of organic compounds having a  $\pi$ -electron conjugated system, and an electrolyte prepared from a material which is chemically stable to the cathode (5) and anode (2) materials and permits the migration of the ion of the anode (2) metal to the cathode (5) for electrochemical reaction with the cathode-active material.



GB 2 068 631 A

2068631

1 / 1

FIG. 1

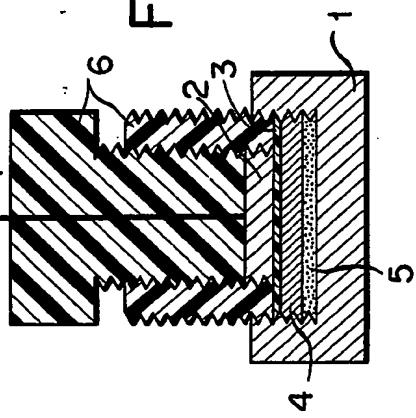
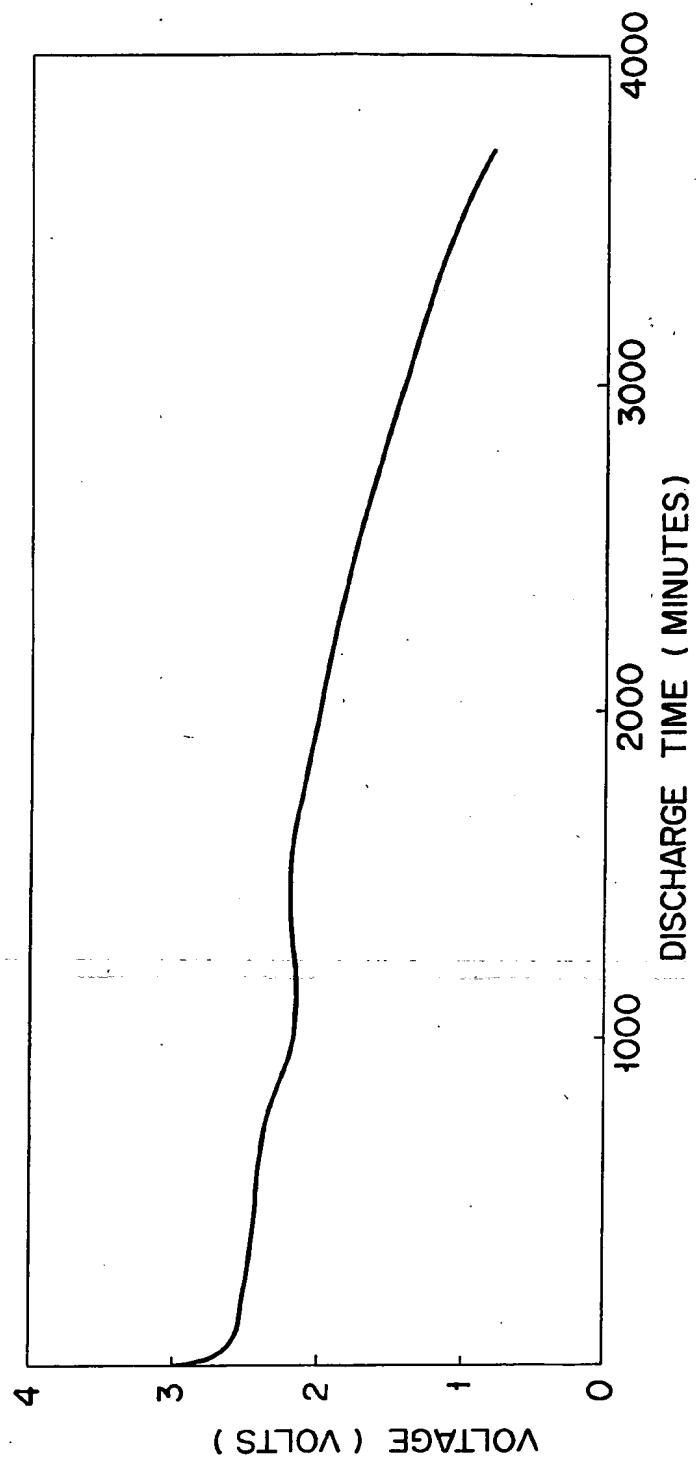


FIG. 2



## SPECIFICATION

## Light weight battery

5 This invention relates to a battery, and more particularly to a battery whose anode contains as the anode-active material a metal selected from the metals of Group Ia, whose cathode contains as the cathode-active material a compound selected from the group consisting of organic compounds having a  $\pi$ -electron conjugated system, and whose electrolyte is formed of a material which does not chemically react with the cathode and anode materials, and permits the migration of ions from the anode material to the cathode for electrochemical reactions with the cathode-active material. 5

10 Hitherto known are numerous types of high energy density battery, in which lithium is used as an anode-active material. A lithium battery is already commercially available in which a cathode-active material is formed of, for example, fluorocarbon,  $Ag_2CrO_4$ ,  $MnO_2$  or  $SOCl_2$ . However, the prior art lithium batteries have the drawback that their capacity and energy density are not always quite satisfactory, and that these 10 batteries are not capable of being recharged. 10

15 Among the secondary batteries using lithium as an anode-active material, a chalcogenide battery (refer to the United States Patent 4,009,052) has the most excellent property in which a cathode-active material is formed of the sulfides, selenides or tellurides of titanium, zirconium, hafnium, niobium, tantalum or vanadium. Among the lithium batteries, the type in which titanium disulfide  $TiS_2$  is used as a cathode-active 15 material was once on the market. Attempts proposed to date which comprise applying an organic compound as a cathode-active material include a process (refer to the specification of United States patent No. 3,532,546) of using liquid ammonia as an electrolyte, and applying, for example, 2,4-dinitroaniline as a cathode-active material; and a process (refer to the Derwent Japanese Patents Report vol. T (1972), No. 72436T) of using liquid ammonia as an electrolyte, and nitroaniline as a cathode-active material. These processes relate to a 20 low temperature-operated battery, because liquid ammonia has a boiling point of  $-33^{\circ}C$ . Nothing is known about the property of a battery based on the application of an electrolyte other than liquid ammonia. Further 20 proposed are the following batteries: 25

(a) United States patent No. Re 27,835 in which a non-aqueous solvent type battery is formed by dissolving  $SO_2$  in electrolyte to use a cathode-active material such as  $N,N,N',N'$ -tetramethyl benzidine which 30 is soluble in a non-aqueous solvent type electrolyte. 30

(b) A lithium battery (United States patent 3,185,590) in which a cathode-active material is formed of a polymer of quinones, for example para-quinone;

(c) A lithium battery (Chemical Abstract 1,270,135 S, 1976, Vol. 85) in which a cathode-active material is formed of 2,3-dichloro-5,6-dicyano-1,4-benzoquinone;

35 (d) A battery (Chemical Abstract 127,012 r, 1976, Vol. 85) in which a cathode-active material is formed of bipyridine; and 35

(e) A battery (United States patent 3,352,720) in which a cathode-active material formed of a complex compound of a halogen such as bromine with the nitrogen atoms of polymeric amine such as polyvinyl pyridine. 40

40 Patents relative to a battery in which Mg or a Mg alloy is used as an anode and an aqueous solution is applied as an electrolyte include a process (the United States patent 2,880,122) of using organic nitro compounds such as nitrobenzotriazole as a cathode active material and a process (the United States patent 2,836,645) of applying 2-hydroxy-1,4-naphthoquinone or phenanthrene quinone as a cathode-active material. This invention has been proposed by way of eliminating the drawbacks of said batteries. 40

45 The primary object of this invention is to provide primary and secondary batteries, whose cathode-active material is composed of an organic compound having a  $\pi$ -electron conjugated system. 45

Another object of the invention is to provide a compact, light weight primary batteries having large energy density.

Still another object of the invention is to provide a secondary batteries which are capable of being 50 discharged and charged over many cycles. 50

The invention is characterized in that with batteries which comprise an anode containing as the anode-active material a metal selected from the Group Ia metals, a cathode, and an electrolyte which does not chemically react with the anode and cathode materials, and permits the migration of the ions from the anode material to the cathode for an electrochemical reaction with the cathode-active material, the cathode 55 material contains as the cathode-active material a compound selected from the group consisting of organic compounds having a  $\pi$ -electron conjugated system. 55

This invention can be more fully understood from the following detailed description when taken in conjunction with the accompanying drawings, in which:

Figure 1 is a vertical sectional view of a battery embodying this invention; and

60 Figure 2 is a curve diagram showing the relationship between a discharge time and a battery voltage of this invention. 60

There will now be described the cathode, anode and electrolyte of a battery according to a preferred embodiment of this invention.

The cathode is prepared from a compound selected from a group of organic compounds having a 65  $\pi$ -electron conjugated system. To describe in greater detail, the cathode should preferably be formed of a 65

compound selected from the following three groups of organic compounds:		
(1) Organic compounds having at least one $-NO_2$ group and $\pi$ -electron conjugated system preferably include		
2,4,7-trinitro-9-fluorenone (TNF)		
5 2,4,5,7-tetranitro-9-fluorenone (TENF)	5	
N-methyl-N-2,4,6-tetranitroaniline (also referred to as nitramine or tetryl)		
o-nitrobenzonitrile		
m-nitrobenzonitrile		
p-nitrobenzonitrile		
10 p-nitrobenzophenone	10	
nitrobenzotriazole		
p-nitrobenzoyl chloride		
m-nitrobenzoyl chloride		
$\gamma$ -(p-nitrobenzyl) pyridine		
15 (2) Organic metal complex compounds having a $\pi$ -electron conjugated system preferably include:	15	
ferric benzoyl acetone [also referred to as tris(1-phenyl-1,3-butanediono) iron]		
bis(8-oxyquinoline) copper (II)		
tris(2-oxy-1,4-naphthoquinone) cobalt (III)		
bis(2-oxy-1,4-naphthoquinone) cobalt (II)		
20 bis(2-oxy-1,4-naphthoquinone) iron (II)	20	
bis(salicyl aldiminate) copper (II)		
tris(salicyl aldiminate) iron (III)		
copper (II) complex of cupferron		
beryllium acetyl acetone		
25 iron (III) acetyl acetone	25	
cobalt (III) acetyl acetone		
cobalt (II) acetyl acetone		
iron (II) naphthoquinone		
(3) Organic compounds having a $\pi$ -electron conjugated system and the ability of forming a complex		
30 compound with a metal preferably include	30	
9,10-phenanthrene quinone		
1,10-phenanthroline		
fluorenone		
$\alpha$ -naphthoquinoline		
35 $\beta$ -naphthoquinoline	35	
2-hydroxy-1,4-naphthoquinone		
cupferron		
The cathode of the battery of this invention is formed by any of the following processes:		
(a) pressing a powder of any of the above-listed cathode-active materials or a mixture of said powder		
40 and a powder of polytetrafluoroethylene binder in the form of a film on a support of, for example, nickel or stainless steel;	40	
(b) mixing a powder of the cathode-active material with an inert electrically conductive powder of, for example, acetylene black, graphite, carbon, or nickel to render to mixed mass electrically conductive, further putting said mixture in a metal vessel with or without a powder of binder such as polytetrafluoroethylene		
45 and compressing the mixed mass; and	45	
(c) mixing the mixture of a powder of the cathode-active material and inert electrically conductive powder in a solution or an emulsion of a binder and spreading the resultant mixture on a support of, for example, nickel or stainless steel in the form of a film, followed by drying.		
One example of this invention is to mix preferably 200% by weight or less of acetylene black based on the		
50 above described cathode-active material and the said cathode-active material with an electrolyte and use the resultant pasty mass as cathode of the battery of the invention. Another example is to mix the cathode-active material with preferably about 10% by weight of a powder of polytetrafluoroethylene binder and preferably 200% by weight or less of acetylene black based on the cathode-active material and press the mixture into a pellet at room temperature for use as the cathode of said battery.	50	
55 Speaking from the characteristic of a primary battery, the cathode-active material is demanded to have a high discharge energy density. From this standpoint, the following compounds which have a high discharge energy density are particularly preferred as cathode-active materials:	55	
2,4,7-trinitro-9-fluorenone		
2,4,5,7-tetranitro-9-fluorenone		
60 N-methyl-N-2,4,6-tetranitroaniline	60	
m-nitrobenzonitrile		
ferric benzoyl acetone		
The compounds listed below are also preferred for the above-mentioned requirement:		
p-nitrobenzonitrile		
65 p-nitrobenzophenone	65	

p-nitrobenzoyl chloride  
 nitrobenzotriazole  
 $\gamma$ -(p-nitrobenzyl) pyridine  
 9,10-phenanthrene quinone  
 5 bis(8-oxyquinoline) copper (II)  
 bis(salicyl aldiminate) copper (II)  
 Speaking of the charge-discharge characteristic of a cathode-active material used as a secondary battery,  
 the following compounds are particularly preferred.  
 ferric benzoyl acetonate  
 10 2,4,7-trinitro-9-fluorenone  
 9,10-phenanthrene quinone  
 N-methyl-N,2,4,6-tetranitroaniline  
 p-nitrobenzoyl chloride  
 There will now be described the anode of a battery embodying this invention. The anode-active material of  
 15 the battery is formed of a metal selected from Group Ia metals, mixtures of said metals themselves or  
 mixtures thereof with any other material. The anode-active material should preferably be formed of lithium.  
 These anode-active materials are used as the anode in the form of a sheet or a sheet pressed on a metal net.  
 There will now be described the electrolyte of the battery of this invention.  
 The electrolyte should be the material which is chemically stable to the cathode and anode materials, and  
 20 permits the migration of the ions of an anode metal to the cathode for electrochemical reaction with the  
 cathode-active material. The electrolyte of the battery of the invention may be used in the form dissolved in a  
 proper solvent, or the solid or molten form. The ion-conducting species of the electrolyte is desired to be the  
 ion of an anode metal.  
 Particularly effective electrolytes are preferably LiC<sub>6</sub>O<sub>4</sub>, LiBF<sub>4</sub>, LiPF<sub>6</sub>, LiAsF<sub>6</sub>, KC<sub>6</sub>O<sub>4</sub>, NaC<sub>6</sub>O<sub>4</sub>, sodium  
 25  $\beta$ -alumina and LiC<sub>6</sub>. Where the electrolyte is used in the form dissolved in a solvent, then said solvent should  
 preferably be ketones, ethers, esters, organic lactones, organic nitriles, organic carbonates, nitrohydrocar-  
 bons and mixtures thereof. The solvent may be formed of, for example, propylene carbonate or a mixture of  
 50 50% of propylene carbonate and 50% of 1,2-dimethoxy ethane.  
 A battery manufactured as described above has a large discharge capacity per unit weight of a  
 30 cathode-active material, providing a light weight battery. The undermentioned example in which  
 2,4,7-trinitro-9-fluorenone is applied as a cathode-active material provides a battery having a discharge  
 capacity about 6, 2 and 8 times larger than MnO<sub>2</sub>, fluoro carbon [(CF)<sub>n</sub>] and TiS<sub>2</sub> respectively used as a  
 cathode-active material in the conventional battery.  
 This invention will be more fully understood by reference to the examples which follow.  
 35 Examples 1 to 16, an organic compound having at least one  $-NO_2$  group and  $\pi$ -electron conjugated  
 system is used as a cathode-active material. In Examples 17 to 24 and 30, an organic metal complex  
 compound having a  $\pi$ -electron conjugated system is used as a cathode-active material. In examples 25 to 30,  
 a compound selected from a group of organic compounds which have a  $\pi$ -electron conjugated system and  
 the ability of forming a complex compound with a metal is used as a cathode-active material. All the  
 40 examples indicate the prominent characteristic of a Li battery embodying this invention.  
 Cathode-active materials applied in the above-mentioned examples and materials used in synthesizing  
 said cathode-active materials were procured from Tokyo Kasei Co., except for ferric benzoyl acetonate and  
 45 2,4,5,7-tetrinitro-9-fluorenone imported from Aldrich Inc. of the United States. Throughout the examples,  
 the discharge capacity and energy density were determined per unit weight of a cathode-active material,  
 unless particularly noted.  
**Example 1**  
 Figure 1 is a vertical sectional view of a battery embodying this invention. Reference numeral 1 denotes a  
 50 nickel-plated brass vessel provided with a circular recess having a diameter of 25 mm. 2 is a circular lithium  
 anode plate having a diameter of 20 mm, 3 a circular porous polypropylene separator having a diameter of  
 55 25 mm; 4 a circular carbon fiber felt having a diameter of 25 mm and a thickness of 3 mm; 5 a cathode  
 mixture; 6 a polytetrafluoroethylene vessel; and 7 a nickel lead wire. The cathode mixture 5 placed in the  
 vessel 6. The lithium anode 2 was set on the compressed mass with the porous polypropylene separator 3  
 60 interposed therebetween. The electrolyte used was formed of a 1 mole/l solution (2 to 5 cc) of LiC<sub>6</sub>O<sub>4</sub>  
 dissolved in propylene carbonate which was distilled and later dehydrated by adding molecular sieves. The  
 porous polypropylene separator 3 and carbon fiber felt 4 were impregnated with the electrolyte thus  
 prepared. The cathode mixture 5 was prepared by mixing 0.05 g of a powder of 2,4,7-trinitro-9-fluorenone  
 65 and 0.05 g of a powder of acetylene black with said electrolyte. A lithium sheet was used as the anode 2 with  
 the porous polypropylene separator 3 disposed between said anode 2 and the carbon fiber felt 4.  
 A battery thus constructed was discharged at a constant current by means of a galvanostat. The discharge  
 voltage of the battery was measured by a high input impedance type voltmeter. An output from the  
 voltmeter was recorded on a pen recorder.  
 65 The battery constructed as described above confined to be discharged 59 hours at a current of 1.57 mA

until the voltage dropped to 1 volt. In this case, the cathode-active material has a discharge capacity of 1,835 ampere-hour (Ah)/kg. The average voltage was 1.60 volts. The energy density was 2,940 watt-hour (Wh)/kg. The open circuit voltage of the battery was 3.15 volts. Figure 2 is a discharge curve diagram, showing variation in the voltage of said battery during the above-mentioned discharge.

5 **Example 2**  
A cathode mixture 5 was prepared by mixing the same type of electrolyte as described above with 0.05 g of a powder of 2,4,7-trinitro-9-fluorenone and 0.05 g of a powder of acetylene black in order to manufacture a battery shown in Figure 1.

10 With a battery manufactured as described above, a constant current discharge was carried out with a discharge current set at different levels as 0.5 mA, 1.57 mA, 3.14 mA, 4.71 mA and 7.80 mA, the results being set forth in Table 1 below. The discharge capacity and energy density given in Table 1 represent values indicated until the battery voltage falls to 1 volt. Since the discharge capacity and energy density reach maximum values at the 1.57 mA discharge, it is preferred to carry out discharge with a discharge current set 15 at 0.5 to 3.14 mA, in case 2,4,7-trinitro-9-fluorenone is used as a cathode-active material.

TABLE 1

20	Discharge current (mA)	Discharge capacity (Ah/kg)	Energy Density (Wh/kg)	20
	0.5	961	2,005	
25	1.57	1,835	2,940	25
	3.14	880	1,300	
	4.71	750	1,350	
30	7.80	260	490	30

**Example 3**

A cathode mixture 5 was prepared by mixing the same type of electrolyte as described above with a powder of 2,4,7-trinitro-9-fluorenone and a powder of acetylene black blended together in the proportions given in Table 2 below in order to manufacture a battery shown in Figure 1. Table 2 below indicates the discharge capacity and energy density of the 2,4,7-trinitro-9-fluorenone which were measured until said battery was discharged at the rate of 1 mA down to 1 volt.

TABLE 2

40	2,4,7-trinitro-9-fluorenone (g)	Acetylene black (g)	Discharge capacity of an active material indicated until the voltage fell to 1 volt (Ah/kg)	Energy density of an active material indicated until the voltage fell to 1 volt (Wh/kg)	Energy density of a cathode mixture indicated until the voltage fell to 1 volt (Wh/kg)	40
45	0.11	0.10	550	1,100	550	45
50	0.05	0.10	750	1,537	512	50
	0.03	0.10	850	1,785	412	

55 The active material given in Table 2 above denotes 2,4,7-trinitro-9-fluorenone. The cathode mixture is formed of 2,4,7-trinitro-9-fluorenone blended with acetylene black. From the energy density per unit amount of the cathode mixture, it is preferred that the 2,4,7-trinitro-9-fluorenone be added at the rate of 0.05 g or more per 0.1 g of acetylene black, namely, 200% by weight or less of the acetylene black based on the 60 2,4,7-trinitro-9-fluorenone be added to the 2,4,7-trinitro-9-fluorenone.

**Example 4**

A battery shown in Figure 1 was manufactured in substantially the same manner as in Example 1, except that 0.015 g of 2,4,7-trinitr -9-fluorenone and 0.025 g of acetylene black were blended with the following electrolyte to provide a cathode mixture. The electrolyte used was formed of a 1 mole/l solution of  $\text{LiC}_2\text{O}_4$  dissolved in a mixture consisting of 1,2-dimethoxy ethane and propylene carbonate mixed in the volumetric ratio of 1:1 which was distilled and later dehydrated.

With the battery thus manufactured, constant current discharge was carried out at 1 mA for 4 hours. Later, constant current charge was carried out at 1 mA for 4 hours each time, indicating a charge discharge capacity of 267 Ah/kg. After the first 4-hour discharge, the voltage of the battery stood at 2.1 volts. After the fifth 10 4-hour discharge, the voltage indicated 1.4 volts. After the tenth 4-hour discharge, the voltage showed 1 volt. After the twentieth 4-hour discharge, the voltage also indicated 1 volt. This event proves that a system allowing for stable charge and discharge was produced in the battery during the 10th and subsequent discharge tests.

**15 Examples 5 to 14**

A battery illustrated in Figure 1 was manufactured in substantially the same manner as in Example 1, except that cathode-active materials having kinds and amounts shown in Table 3 below and acetylene black having amounts indicated therein were blended with the same type of electrolyte as used in Example 1 to provide a cathode mixture.

20 With the battery thus manufactured, constant current discharge was carried out at 1 mA, until the battery voltage dropped to 1 volt (Examples 5 to 8), and also until the battery voltage fell to 2 volts (Examples 9 to 14). Table 3 below shows the discharge capacity and energy density which were indicated until the battery voltage was reduced to 1 volt (Examples 5 to 8) and also to 2 volts (Examples 9 to 14).

TABLE 3

Exam- ple	Cathode-active material Compound	Amount (g)	Acetylene Black Amount (g)	OCV* (V)	Discharge capacity (Ah/kg)	Energy density (Wh/kg)
5	2,4,5,7-tetranitro-9-fluorenone	0.03	0.1	3.35	958	1,913
6	N-methyl-N-(2,4,6-tetranitroaniline)	0.1	0.1	3.04	930	1,994
7	p-nitrobenzoyl chloride	0.05	0.05	3.33	530	1,290
8	m-nitrobenzoyl chloride	0.05	0.05	3.10	400	890
9	o-nitrobenzonitrile	0.05	0.05	2.40	240	536
10	m-nitrobenzonitrile	0.05	0.05	3.10	740	1,610
11	p-nitrobenzonitrile	0.05	0.05	3.05	500	1,116
12	p-nitrobenzophenone	0.05	0.05	3.05	430	967
13	$\gamma$ -(p-nitrobenzyl) pyridine	0.05	0.05	3.15	400	1,093
14	5-nitrobenzotriazole	0.05	0.05	3.05	450	895

\* Open circuit voltage.

*Example 15*

A battery shown in Figure 1 was manufactured in substantially the same manner as in Example 1, except that 0.04 g of N-methyl-N,2,4,6-tetranitroaniline and 0.05 g of acetylene black were mixed with the same type of electrolyte as used in Example 1 to provide a cathode mixture.

5 With the battery thus manufactured, constant current discharge was carried out at 1 mA for 4 hours. Later, constant current charge was carried out at 1 mA for 4 hours, indicating a charge discharge capacity of 100 Ah/kg. This cycle of discharge and charge was repeated. After the first 4-hour discharge, the battery voltage stood at 2.60 volts. After the fifth 4-hour discharge, the battery voltage indicated 2.40 volts. After the tenth 4-hour discharge, the battery voltage showed 2.25 volts.

10 10

*Example 16*

A battery shown in Figure 1 was manufactured in substantially the same manner as in Example 1, except that 0.025 g of p-nitrobenzoyl chloride and 0.04 g of acetylene black were blended with the same type of electrolyte as used in Example 1 to provide a cathode mixture. With the battery thus manufactured, constant 15 current discharge was carried out at 1 mA for 5 hours, indicating a discharge capacity of 200 Ah/kg. Later, constant current charge was carried out at 1 mA for 5 hours, indicating a charge capacity of 200 Ah/kg. This cycle of discharge and charge was repeated. At the first cycle test, a discharge capacity until 1 volt stood at 192 Ah/kg. At the seventh cycle test, a discharge capacity until 1 volt indicated 180 Ah/kg. At the tenth cycle test, a discharge capacity until 1 volt showed 180 Ah/kg. At the fourteenth cycle test, a discharge capacity 20 until 1 volt indicated 160 Ah/kg.

20 20

*Examples 17 to 20*

A battery shown in Figure 1 was manufactured in substantially the same manner as in Example 1, except that cathode-active materials having kinds and amounts given in Table 4 below and acetylene black having 25 amounts shown therein were blended with the same type of electrolyte as used in Example 1 to provide a cathode mixture. The open circuit voltage of the battery is also set forth in Table 4.

With the battery thus manufactured, constant current discharge was carried out at 1 mA, until the battery voltage dropped to 1 volt. Table 4 also indicates the discharge capacity and energy density indicated until the battery voltage fell to 1 volt.

30

TABLE 4

Exam- ple	Cathode-active material Compound	Acetylene Black		OCV <sup>(1)</sup> (V)	Discharge capacity (Ah/kg)	Energy density (Wh/kg)
		Amount (g)	Amount (g)			
17	ferric benzoyl acetonate	0.05	0.05	3.15	910	1,367
18	bis(8-oxy- quinoline) copper (II)	(2)	0.05	0.1	3.12	1,000
19	tris(2-oxy-1,4- naphthoquinone) cobalt (III)	(3)	0.05	0.05	3.08	610
20	bis(salicyl aldiminate) copper (II)	(4)	0.05	0.05	2.90	650
						1,540

Note: (1) Open circuit voltage.

(2) Obtained by adding copper acetate to an alcohol solution of 8-hydroxyquinoline, heating the mass for sometimes on a water bath, separating a precipitate produced in the alcohol solution by a filtration and recrystallizing the precipitate by using a hot alcohol.

(3) Synthesized as in Example 18 from 2-hydroxy-1,4-naphthoquinone and cobalt acetate.

(4) Obtained in the form of a precipitate after the process of adding ammonia water and cobalt (II) acetate to an alcohol solution of salicyl aldehyde, and thereafter heating the mass.

**Example 21**

0.05 g of ferric benzoyl acetonate powder and 0.05 g of acetylene black were blended with the same type of electrolyte as used in Example 1 to provide a cathode mixture, thereby manufacturing a battery shown in Figure 1. With the battery thus manufactured, constant current discharge was carried out, with a discharge current set at 1 mA, 1.57 mA, 3.14 mA and 4.71 mA, the results of the tests being set forth in Table 5 below.

5

TABLE 5

	Discharge current (mA)	Discharge capacity (Ah/kg)	Energy density (Wh/kg)	
10	1	910	1,370	10
15	1.57	570	760	15
	3.14	150	210	
	4.71	180	260	

20 The discharge capacity and energy density indicated in Table 5 above represent values indicated until the battery voltage dropped to 1 volt. As seen from Table 5, the discharge capacity and energy density indicate larger values at smaller discharge currents. Where, therefore, ferric benzoyl acetonate is applied as a cathode-active material, it is preferred to carry out discharge at a smaller discharge current than 1.57 mA.

25 **Example 22**  
A battery shown in Figure 1 was manufactured in substantially the same manner as in Example 1, except that 0.03 g of ferric benzoyl acetonate and 0.03 g of acetylene black were blended with the same type of electrolyte as used in Example 1 to provide a cathode mixture. With the battery thus manufactured, constant 30 current discharge was carried out at 1 mA for 1.5 hours, indicating a discharge capacity of 50 Ah/kg. Later, constant current charge at 1 mA for 1.5 hours was carried out, indicating a charge capacity of 50 Ah/kg. The above-mentioned cycle of discharge and charge was repeated 120 times.

Since the discharge capacity of a cathode-active material indicated until the battery voltage dropped to 1 volt stood at 910 Ah/kg, the above-mentioned cycle of charge and discharge correspond to a 5.5% capacity. 35 At the first discharge test, a minimum discharge voltage indicated 1.51 volts. At the fiftieth discharge test, a minimum discharge voltage showed 1.86 volts. At the ninetieth discharge test, a minimum discharge voltage indicated 1.78 volts. At the last discharge test, a minimum discharge voltage gave 1.41 volts.

40 **Example 23**  
A battery shown in Figure 1 was manufactured in substantially the same manner as in Example 1, except that 0.01 g of ferric benzoyl acetonate and 0.03 g of acetylene black were blended with the same type of electrolyte as used in Example 1 to provide a cathode mixture.

45 With the battery thus manufactured, constant current discharge was carried out at 1 mA for 5 hours, indicating a discharge capacity of 500 Ah/kg. Later, constant current charge was carried out at 1 mA for 5 hours, indicating a charge capacity of 500 Ah/kg. The above-mentioned cycle of discharge and charge was repeated 98 times.

50 Since the discharge capacity of a cathode-active material indicated until the battery voltage dropped to 1 volt was 910 Ah/kg, the above-mentioned cycle of charge and discharge correspond to a 55% capacity. At the first discharge test, a minimum discharge voltage was 1.31 volts. At the fortieth discharge test, a minimum discharge voltage was 1.09 volts. At the eightieth discharge test, a minimum discharge voltage was 1.08 volts. At the last discharge test, a minimum discharge voltage was 1.20 volts.

55 **Example 24**  
A battery shown in Figure 1 was manufactured in substantially the same manner as in Example 1, except that 0.03g of tris(2-oxy-1,4-naphthoquinone) cobalt (III) and 0.03 g of acetylene black were mixed with the same type of electrolyte as used in Example 1 to provide a cathode mixture.

60 With the battery thus manufactured, constant current discharge was carried out at 1 mA for 6 hours, indicating a discharge capacity of 200 Ah/kg. After a pause of 1 hour, constant current charge was carried out at 1 mA for 6 hours, indicating a charge capacity of 200 Ah/kg. The charge was stopped for 1 hour. This cycle of discharge and charge was repeated. After the first 6-hour discharge test, the battery voltage stood at 0.80 volt. After the third 6-hour discharge test, the battery voltage indicated 1.10 volts. After the fourth 6-hour discharge test, the battery voltage showed 1.20 volts. After the sixth 6-hour discharge test, the battery voltage gave 0.95 volts.

**Examples 25 and 26**

A battery shown in Figure 1 was manufactured in substantially the same manner as in Example 1, except that cathode-active materials having kinds and amounts shown in Table 6 below and acetylene black having amounts indicated therein were blended with the same type of electrolyte as used in Example 1 to provide a cathode mixture. Table 6 below also shows the open circuit voltage of the battery thus manufactured. 5

With said battery, constant current discharge was carried out at 1 mA, until the battery voltage fell to 1 volt. The discharge capacity and energy density indicated until the battery voltage was reduced to 1 volt are set forth in Table 6 below.

		TABLE 6						
10	15	Cathode-active material	Acetylene Black	OCV* (V)	Discharge capacity (Ah/kg)	Energy density (Wh/kg)	10	
Example	Compound	Amount (g)	Amount (g)				15	
25	9,10-phenanthrene quinone	0.03	0.1	2.82	520	1,130		
20	2-hydroxy-1,4-naphthoquinone	0.05	0.05	2.90	410	801	20	

\* Open circuit voltage.

25	<b>Example 27</b>	25
A cathode mixture was prepared by mixing the same type of electrolyte as used in Example 1 with 9,10-phenanthrene quinone powder and acetylene black blended together in the proportions given in Table 7 below in order to manufacture a battery shown in Figure 1.		
30	Table 7 sets forth the discharge capacity and energy density of the 9,10-phenanthrene quinone indicated until the battery voltage decreased to 1 volt. The 9,10-phenanthrene quinone given in Table 7 denotes a cathode-active material. As determined from the energy density per unit weight of the cathode mixture (9,10-phenanthrene quinone and acetylene black), it is preferred that more than 0.1 g of said 9,10-phenanthrene quinone be mixed with 0.1 g of acetylene black, namely, less than 200% by weight of acetylene black be added to the 9,10-phenanthrene quinone.	30
35		35

TABLE 7

40	9,10-phenanthrene quinone (g)	Acetylene black (g)	Discharge capacity of the cathode-active material indicated until the battery voltage dropped to 1 volt (Ah/kg)	Energy density of the cathode-active material indicated until the battery voltage fell to 1 volt (Wh/kg)	Energy density of the cathode mixture indicated until the battery voltage was reduced to 1 volt (Wh/kg)	40
45	0.1	0.1	320	710	355	
50	0.03	0.06	350	773	257	50
	0.03	0.1	519	1,130	261	

**Example 28**

55	A battery shown in Figure 1 was manufactured in substantially the same manner as in Example 1, except that 0.04 g of 9,10-phenanthrene quinone and 0.10 g of acetylene black were blended with the same type of electrolyte as used in Example 1 to provide a cathode mixture. With the battery thus manufactured, constant current discharge was carried out at 1 mA for 7 hours. Later, constant current charge was carried out at 1 mA for 7 hours, indicating a charge discharge capacity of 233 Ah/kg. This cycle of discharge and charge was	55
60	repeated. The discharge capacity until 2 volts at the first discharge test indicated 175 Ah/kg. The discharge capacity until 2 volts at the fourth discharge test was 133 Ah/kg. The discharge capacity until 2 volts at the fifth discharge test was 125 Ah/kg.	60

**Example 29**

A battery shown in Figure 1 was manufactured in substantially the same manner as in Example 1, except that 0.03 g of 2-hydroxy-1,4-naphthoquinone and 0.03 g of acetylene black were blended with the same type of electrolyte as used in Example 1 to provide a cathode mixture.

5 With the battery thus manufactured, constant current discharge was carried out at 1 mA for 6 hours, indicating a discharge capacity of 200 Ah/kg. After a pause of 1 hour, constant current charge was carried out at 1 mA for 6 hours, indicating a charge capacity of 200 Ah/kg. The charge was stopped for 1 hour. This cycle of discharge and charge was repeated. After the 6-hour discharge of the first test cycle, the battery voltage stood at 2.20 volts. After the 6-hour discharge of the fourth test cycle, the battery voltage indicated 1.68 volts.

10 After the 6-hour discharge of the eighth test cycle, the battery voltage showed 1.10 volts. After the 6-hour discharge of the twenty-first test cycle, the battery voltage was 0.68 volt.

**Example 30**

A battery shown in Figure 1 was manufactured in substantially the same manner as in Example 1, except that 0.1 g of each of the cathode-active materials listed in Table 8 below and 0.1 g of acetylene black were blended with the same type of electrolyte as used in Example 1 to provide a cathode mixture.

With the batteries thus manufactured, constant current discharge was carried out at 1 mA, until the battery voltage dropped to 1 volt. The tested batteries indicated the open circuit voltage, the discharge capacity, energy density and average voltage listed in Table 8 below.

20

TABLE 8

	Cathode-active material	OCV* (V)	Discharge capacity (Ah/kg)	Energy density (Wh/kg)	Average voltage (V)	
25	Bis(2-oxy-1,4-naphthoquinone) copper (II)	3.25	130	244	1.88	25
30	Bis(2-oxy-1,4-naphthoquinone) iron (II)	2.35	280	484	1.79	30
35	Tris (salicyl aldiminate) iron (III)	3.40	455	665	1.46	35
40	Copper (II) complex of cupferron	3.03	310	441	1.42	40
45	Beryllium acetyl acetonate	3.15	130	153	1.18	45
50	Iron (III) acetyl acetonate	2.68	235	353	1.50	50
55	Cobalt (III) acetyl acetonate	3.25	315	495	1.57	55
60	Cobalt (II) acetyl acetonate	2.90	430	648	1.51	60
65	$\alpha$ -naphthoquinoline	3.10	112	155	1.38	65
	$\beta$ -naphthoquinoline	2.40	130	187	1.44	
	Cupferron	2.95	420	585	1.39	
	1,10-phenanthroline	2.90	450	640	1.43	
	Fluorenone	2.95	165	267	1.62	
	* Open circuit voltage.					

## CLAIMS

1. A light weight battery which comprises:  
an anode containing one metal selected from the Group Ia metals;
- 5 a cathode containing one compound selected from the group consisting of organic compounds having a  $\pi$ -electron conjugated system; and  
an electrolyte which does not chemically react with the anode and cathode materials, and permits the migration of the ions of the anode material to the cathode for electrochemical reaction with the cathode-active material.
- 10 2. The battery according to claim 1, wherein the cathode contains one compound selected from the group consisting of organic compounds having at least one  $-NO_2$  group and a  $\pi$ -electron conjugated system. 10
3. The battery according to claim 1, wherein the cathode contains one compound selected from the group consisting of organic metal complex compounds having a  $\pi$ -electron conjugated system.
- 15 4. The battery according to claim 1, wherein the cathode contains one compound selected from the group consisting of organic compounds having a  $\pi$ -electron conjugated system and the ability of forming a complex compound with a metal. 15
5. The battery according to claim 2, wherein the cathode contains one compound selected from the group consisting of 2,4,7-trinitro-9-fluorenone; 2,4,5,7-tetranitro-9-fluorenone; N-methyl-N,2,4,6-tetranitroaniline; o-nitrobenzonitrile; m-nitrobenzonitrile; p-nitrobenzonitrile; p-nitrobenzophenone; nitrobenzotriazole; p-nitrobenzoyl chloride; m-nitrobenzoyl chloride; and  $\gamma$ -(p-nitrobenzyl) pyridine. 20
6. The battery according to claim 3, wherein the cathode contains one compound selected from the group consisting of ferric benzoyl acetone; bis(8-oxyquinoline) copper (II); tris(2-oxy-1,4-naphthoquinone) cobalt (III); bis(2-oxy-1,4-naphthoquinone) iron (II); bis(salicyl aldiminate) copper (II); tris (salicyl aldiminate) iron (III); copper (II) complex of cupferron; beryllium acetyl acetone; iron (III) acetyl acetone; cobalt (III) acetyl acetone; cobalt (II) acetyl acetone; and iron (II) naphthoquinonate. 25
7. The battery according to claim 4, wherein the cathode contains one compound selected from the group consisting of 9,10-phenanthrene quinone; 1,10-phenanthroline; fluorenone;  $\alpha$ -naphthoquinoline;  $\beta$ -naphthoquinoline; 2-hydroxy-1,4-naphthoquinone; and cupferron. 30
8. The battery according to claim 1, wherein the anode contains lithium.
9. The battery according to claim 1, wherein the electrolyte is a solution prepared by dissolving  $LiC_6O_4$  in propylene carbonate.
10. The battery according to claim 2, wherein the cathode contains a mixture of 2,4,7-trinitro-9-fluorenone and 200% by weight or less of a powder of acetylene black based on said 2,4,7-trinitro-9-fluorenone; the anode contains lithium metal; and the electrolyte is prepared by dissolving  $LiC_6O_4$  in propylene carbonate or a mixture of propylene carbonate and 1,2-dimethoxy ethane. 35
11. The battery according to claim 2, wherein the cathode contains a mixture of 2,4,5,7-tetranitro-9-fluorenone and 200% by weight or less of a powder of acetylene black based on said 2,4,5,7-tetranitro-9-fluorenone; the anode contains lithium metal; and the electrolyte is prepared by dissolving  $LiC_6O_4$  in propylene carbonate or a mixture of propylene carbonate and 1,2-dimethoxy ethane. 40
12. The battery according to claim 2, wherein the cathode contains a mixture of N-methyl-N,2,4,6-tetranitroaniline and 200% by weight or less of a powder of acetylene black based on said N-methyl-N,2,4,6-tetranitroaniline; the anode contains lithium metal; and the electrolyte is prepared by dissolving  $LiC_6O_4$  in propylene carbonate or a mixture of propylene carbonate and 1,2-dimethoxy ethane. 45
13. The battery according to claim 2, wherein the cathode contains a mixture of p-nitrobenzoyl chloride and 200% by weight or less of a powder of acetylene black based on said p-nitrobenzoyl chloride; the anode contains lithium metal; and the electrolyte is prepared by dissolving  $LiC_6O_4$  in propylene carbonate or a mixture of propylene carbonate and 1,2-dimethoxy ethane.
- 50 14. The battery according to claim 3, wherein the cathode contains a mixture of ferric benzoyl acetone and 200% by weight or less of a powder of acetylene black based on said ferric benzoyl acetone; the anode contains lithium metal; and the electrolyte is prepared by dissolving  $LiC_6O_4$  in propylene carbonate or a mixture of propylene carbonate and 1,2-dimethoxy ethane. 50
15. The battery according to claim 4, wherein the cathode contains a mixture of 9,10-phenanthrene quinone and 200% by weight or less of a powder of acetylene black based on said 9,10-phenanthrene quinone; the anode contains lithium metal; and the electrolyte is prepared by dissolving  $LiC_6O_4$  in propylene carbonate or a mixture of propylene carbonate and 1,2-dimethoxy ethane. 55
16. A light weight battery, substantially as hereinbefore described with reference to the accompanying drawings and Examples.